

## Physical Chemistry Engel 3rd Edition Solution Manual Pdfrar [BEST]

Objective: The main objective of this Chapter is to extend our previous results and to provide the formulas for an alternative and general way of obtaining the  $E^{\circ}$  of several solutions. Subjects: The subject for this lecture are the following chemical reaction Equations 1-5. In Equation 1, there is one mole of  $H_3O^+$  and six mole of  $H_2CO_2$ . Equation 2 shows that one mole of  $H_2CO_2$  reacts with three moles of  $H_3O^+$  to yield four mole of  $HCO_3^-$ , one mole of  $H^+$  and one mole of  $H_2O$ . In Equation 3, one mole of  $KIO_3$  reacts with one mole of  $H_2O$  to yield one mole of  $KIO_4$  and one mole of  $H^+$  and one mole of  $H_2O$ . Equation 4 shows that two moles of  $KIO_4$  reacts with one mole of  $H_2O$  to yield two mole of  $KH_2PO_4$ . Equation 5 shows that two moles of  $KH_2PO_4$  reacts with two moles of  $H^+$  to yield one mole of  $KOH$  and one mole of  $H_2O$ . Observe that, this Chapter has three objectives. The first objective of this chapter is to show that the  $E^{\circ}$  of  $K^+$  carbonate solution is 8.25 at STP. Secondly, the object of this chapter is to provide the formulas for a general method for calculating the  $E^{\circ}$  of various solutions. The approach is similar to that used for determining the  $E^{\circ}$  of  $Na^+$  carbonate solution. In this approach, the  $E^{\circ}$  is first calculated by (1), then the  $E^{\circ}$  of the solution containing  $K^+$  is obtained by calculating  $E^{\circ}$  of  $KOH$  solution, as is given in (2). Finally, the  $E^{\circ}$  of the solution containing  $K^+CO_3$  solution is obtained by calculating the  $E^{\circ}$  of  $KH_2PO_4$ , as is given in (3). In this lecture, we will therefore deal with the following chemical reaction Equations 6-9. Observe that, all the data required for Equation 6 are given in the above chapter. In Equation 6, the  $E^{\circ}$  of  $H_2CO_3$  solution is known. Also, when  $K^+CO_3$  solution is formed, one mole of  $H_2CO_3$  and two moles of  $CO_2$  will be present. Equ

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